Honors Chemistry Daily EQs

Unit # \_\_8\_\_\_\_

Dates \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_



1. Answer the questions each day during the first 10 minutes of class.
2. If you are absent, you must make-up the missed questions the day you return.

**Score**

1. Write the correct answers for any questions you missed.
2. Turn in this assignment with your unit pack at the end of each unit.

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|  | Daily Warm-Ups: |
| Unit 8 Day 1 | Write balanced equations for the following reaction:   1. Sodium oxide is added to water 2. barium nitrate solution reacts with lithium metal   A calcium chloride hydrate has a mass of 4.72 g. After heating for several minutes the mass of the anhydrate is found to be 3.56 g. Use this information to determine the formula for the hydrate.  Convert 25.0 g of chlorine gas to L at STP. |
| Unit 8 Day 2 | If magnesium oxide is reacted with potassium:   1. What is the balanced chemical equation? 2. If you start with 5.0 moles of magnesium oxide how many moles of Mg are formed? 3. If you start with 5.0 moles of magnesium oxide how many grams of potassium oxide are formed?   A solution containing 3.50 g of sodium phosphate is mixed with a solution containing excess magnesium nitrate.   1. Write a balanced equation 2. How many grams of magnesium phosphate can be produced? |
| Unit 8 Day 3 | Given: 2 H2 (g) + O2 (g) 🡪 2 H2O (l)   1. How many grams of water can be produced from 5 g of oxygen gas? 2. How many grams of water can be produced from 5 L of hydrogen gas? (at STP) 3. What is your limiting reagent? 4. If you produce 1.10 g of water from 5 L of hydrogen gas what is your percent yield? |
| Unit 8 Day 4 | Given: 3 H2SO4 + 2 Al(OH)3 🡪 Al2(SO4)3 + 6 H2O   1. What is the mole ratio between Al(OH)3 and H2O? 2. How many moles of water are produced from 3 moles of H2SO4? 3. How many grams of water can be produced from 2 moles of Al(OH)3? 4. If you have 3.5 g of aluminum hydroxide how many moles of aluminum sulfate can be produced? |
| Unit 8 Day 5 | Then 5.0 g of aluminum reacts with 5.0 g of chlorine gas to produce aluminum chloride according to the reaction:    **2 Al + 3 Cl2 🡪 2 AlCl3**   1. How many grams of AlCl3 are produced? 2. What is the limiting reactant? 3. What is the excess reactant? How much of the excess reactant remains? |