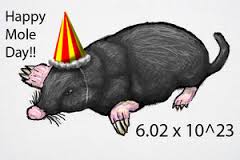
Honors Chemistry Daily EQs



Unit # 6



1. Answer the questions each day during the first 10 minutes of class.
2. If you are absent, you must make-up the missed questions the day you return.

**Score**

1. Write the correct answers for any questions you missed.
2. Turn in this assignment with your unit pack at the end of each unit.

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|  | Daily Warm-Ups: |
| Unit 6 Day 1 | **Copy** and **fill** in the following table:   |  |  |  |  | | --- | --- | --- | --- | | **Element/ion** | **# of protons** | **# of neutrons** | **# of electrons** | | Cr |  |  |  | | Mg+2 |  |  |  | |  | 33 |  | 36 | | N-3 |  |  |  | | Uranium-235 |  |  |  | |
| Unit 6 Day 2 | * Convert the following:   A. 2.2 g of Al to moles B. 2.2 x10 22 molecules water to moles  C. 55 g of C to atoms   * Find the molar mass of the following:   A. BaCl2 B. Cu(NO2)2 C. aluminum sulfide  Image result for mole day |
| Unit 6 Day 3 | Given 35.1 g of Carbon dioxide gas (CO2) convert to   1. Moles B. molecules C. liters (at STP)   What is the percentage of **sulfur** in calcium sulfate (CaSO4)? |
| Unit 6 Day 4 | Calculate the empirical formula of a compound that contains 1.67 g of Cerium, Ce and 4.54 g of iodine, I .  A compound is 53.3 % O and 46.7 % Si solve for the following information:  A. find its empirical formula B. If its molar mass is 180 g/mol what is it’s molecular formula? |
| Unit 6 Day 5 | Image result for mole day  Find the formula of a hydrate that is composed of 48.86 % Magnesium sulfate and 51.14 % water. |